

Viziscience®

ALL-IN-ONE AP Chemistry Resource



VIZ!SCIENCE



Hello there! My name is Sarah and I'm a 4th year AP chemistry teacher.



I used to celebrate decent pass rates and feel confident in my teaching abilities.



Then the pandemic hit, and the year of remote learning left my students lacking in essential foundation knowledge.



My students struggled even with basics such as stoichiometry. I'm torn between filling these gaps and covering the demanding AP Chemistry curriculum.



Despite incorporating textbooks, worksheets, videos, and materials from numerous sources, my students are falling behind. I'm not sure what else I can do.



You're not alone! Many teachers are feeling burn out teaching this difficult class.



We have a solution that's helped countless dedicated teachers like you, and we can't wait to share it with you.



Cool! I can't wait to hear about it!

Welcome to Viziscience®

Today, we're thrilled to show you how our platform can transform the classroom for you. Imagine having everything you need – interactive lessons, engaging concept videos, practical simulations, labs, and quizzes – all in one place, easy to access and use.

One of the unique aspects of our platform is how it provides a cohesive and sequential learning experience which is hard to achieve using a mix and match of materials from various sources. We understand that teachers often struggle with their lessons feeling disjointed.

The AP Chemistry curriculum is challenging, and keeping students on track can be tough. Viziscience® is meticulously designed to guide students through the curriculum step by step, ensuring they build on each concept methodically before moving to the next.

What's really exciting is the flexibility Viziscience® offers. You, the educators, can tailor the learning path to fit the unique needs of each of your students. Whether it's revisiting a tricky concept or pushing ahead, you have the control. Students facing difficulties can independently catch up without slowing down the rest of the class

So, let's dive in and see how Viziscience® not only supports but enhances the learning experience, ensuring your students are not just keeping up, but excelling in AP Chemistry.





Viziscience

The Viziscience AP chemistry course is made up of 4 sections to hone your students' skills:

Interactive Modules

Interactive Labs

Multiple Choice Practice

Summer Prep

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Course categories: Advanced Chemistry



Viziscience® Interactive Modules

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Viziscience® Interactive Labs

ENTER THIS COURSE

Multiple Choice Practice Questions

ENTER THIS COURSE

AP Chem Summer Prep (2023/24)

ENTER THIS COURSE

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Welcome to Viziscience Interactive Modules

Let's explore how Viziscience® aligns with the AP Chemistry Exam Description (CED), ensuring that your students fully benefit from their learning experience.

The AP Chemistry CED is structured around four core ideas, commonly referred to as the 'Big Ideas'. These Big Ideas form the foundation of the course and are critical in shaping students' understanding of chemistry. To bring these ideas to life, the curriculum is further divided into nine main units, each delving into specific topics and concepts.

In Viziscience®, we've mirrored this structure to provide a seamless and intuitive learning path.

Math Review

Important Math Skills

- Exponents and Scientific Notation
- Significant Figures - Quick Review
- Calculator & Exponents
- Logarithm & pH
- Mental Math - Tips and Tricks

UNIT 1

1. Atomic Structure and Properties

- 1.1 Atoms, Isotopes, Atomic Weight
- 1.2 Mole & Molar Mass
- 1.3 The Periodic Table
- 1.4 (b) Composition
- 1.5 (i) Naming
- 1.5 (ii) Na
- 1.5 (iii)
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- 2.00

Big Ideas

The big ideas spiral across units.

	Unit 1	Unit 2	Unit 3
Big Idea 1	Atomic Structure and Properties	Molecular and Ionic Compound Structure and Properties	Intermolecular Forces and Properties
Big Idea 2			
Big Idea 3			



UNIT 2

2. Molecular and Ionic Compound Structure and Properties

- 2.1 Lewis
- 2.2 Lewis
- 2.3 Formal
- Lewis
- 2.4 Ionic Solids, M
- 2.5 VSEPR (PhET)
- 2.6 Bond Hybridization



UNIT 3

3. Intermolecular Forces and Properties

- 3.1 (i) Intermolecular Forces (IMF)
- 3.1 (ii) IMF, Solubility & Visual Representations
- Intermolecular Forces
- 3.2 Avogadro's Law & Dalton's Law
- 3.3 Kinetic Molecular Theory
- 3.4 Ideal Gas Law & Deviation
- 3.5 Electromagnetic Radiation
- 3.6 Beer-Lambert Law

UNIT 4

4. Chemical Reactions

- 4.1 Balancing Equations
- 4.2 Balancing Complex Equations
- 4.3 Stoichiometry
- 4.4 Limiting Reactant & Percent Yield

Viziscience complements CED topics with lessons and additional subtopics designed to support foundational learning.

Let's take the example of teaching Unit 1.1 Moles and Molar Mass. To prepare your students for this topic, assign Viziscience's additional subtopic VZ1.1 'Atoms and Isotopes' as prerequisites to the course.

VZ1.1 equips students with essential knowledge about what contributes to an atom's mass and how atomic mass is calculated through the relative abundance of isotopes. Furthermore, it offers a comprehensive review of isotopes and delves into the practical application of Mass Spectrometry.

If your students lack prior knowledge of this topic, our materials are really helpful. Even for students who have previously covered this subject, it's an excellent opportunity for concept review. This foundation equips your students to make the most of your class as they come well prepared.

Interactive Modules

To help teachers in navigating the CED alignment, we provide an easy to use color-coded roadmap.

This resource will guide teachers in effectively utilizing our supplementary units to bridge any gaps as needed.

CED		Viziscience	
CED Unit 1 Atomic Structure and Properties	Recommended Viziscience Topics	Viziscience Unit 1	
1.1 Moles and Molar Mass		1.1 Atoms, Isotopes, Atomic Weight	
1.2 Mass Spectroscopy of Elements		1.2 Mole & Molar Mass	
1.3 Elemental Composition of Pure Substances		1.3 The Periodic Table	
1.4 Composition of Mixtures		1.4 (a) Elemental Composition of Pure Substances	
1.5 Atomic Structure and Electron Configuration		1.4 (b) Composition of Mixtures	
1.6 Photoelectron Spectroscopy	Lab 2	1.5 (i) Naming Compounds (Covalent)	
1.7 Periodic Trends		1.5 (ii) Naming Compounds (Ionic)	
1.8 Valence Electrons and Ionic Compounds		1.5 (iii) Naming Polyatomic Ions	
		1.6 Electron Configurations	
		1.7 The Order of Electron Filling	
		1.8 Photoelectron Spectroscopy (PES)	
		1.9 Periodic Trends	

On the left, you'll find the CED order of topics, while the right table provides the recommended order as suggested by Viziscience. You have the flexibility to use our recommended order, align the topics with the CED order, or adapt them to fit your own teaching approach.



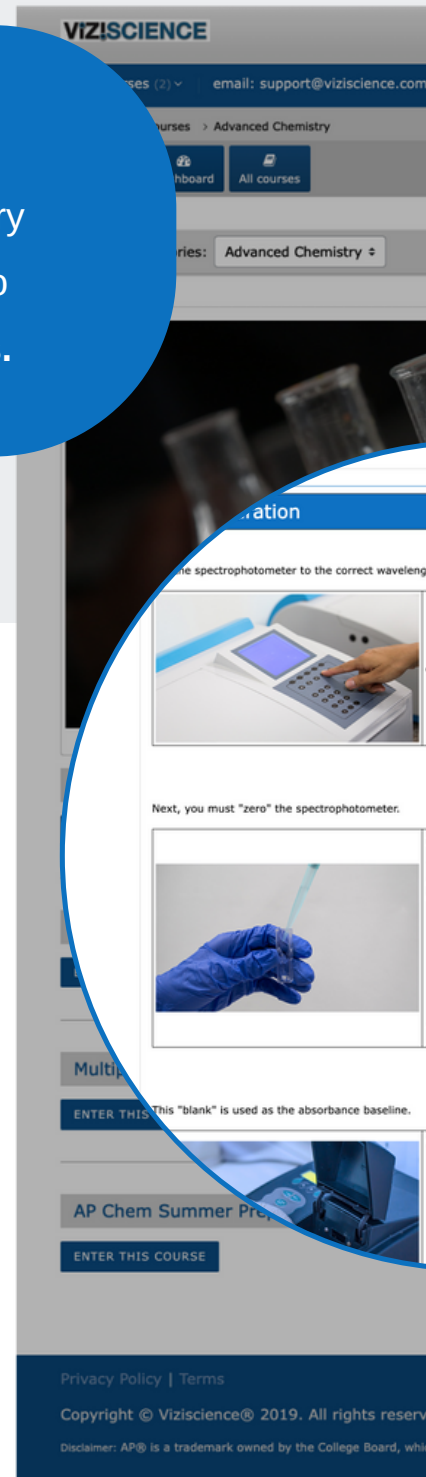
LABS

Laboratory experiences are integral to the AP Chemistry curriculum, offering students invaluable opportunities to apply theoretical knowledge to real-world experiments.



BRIDGING THE GAP WITH VIRTUAL LABS

When it's not possible to conduct traditional wet labs, Viziscience offers a practical and effective solution to bring laboratory experiences into your lessons. Our virtual labs are designed to closely **simulate the real laboratory** experience. Students start by grasping the theory before moving on to learn how to set up equipment, perform experiments, and carefully gather and analyze data. **We put a strong focus on developing skills essential for exams, such as the ability to spot errors and understand their impact on experimental results.**





Interactive Labs

Here is the list of Viziscience's interactive and engaging labs to support the AP chemistry curriculum.

Viziscience
Labs
1. Ionic & Covalent Compound Lab
2. Flame Test Lab
3. Paper Chromatography
4. Molar Mass of a Gas
5. Thin Layer Chromatography
6. Single Displacement
7. Gravimetric Analysis
8. Strong Acid-Strong Base Titration
9. Kinetics Theory
10. Calorimetry Lab
12. Redox Titration
Equilibrium Labs
13. Le Chatelier's Principle & The Haber Process
14. KEQ LabPage
a) KEQ: Experiment & data
b) KEQ: Calculations & Post Lab
15. Solubility & KSP

Multiple Choice Practice

Multiple Choice Quizzes

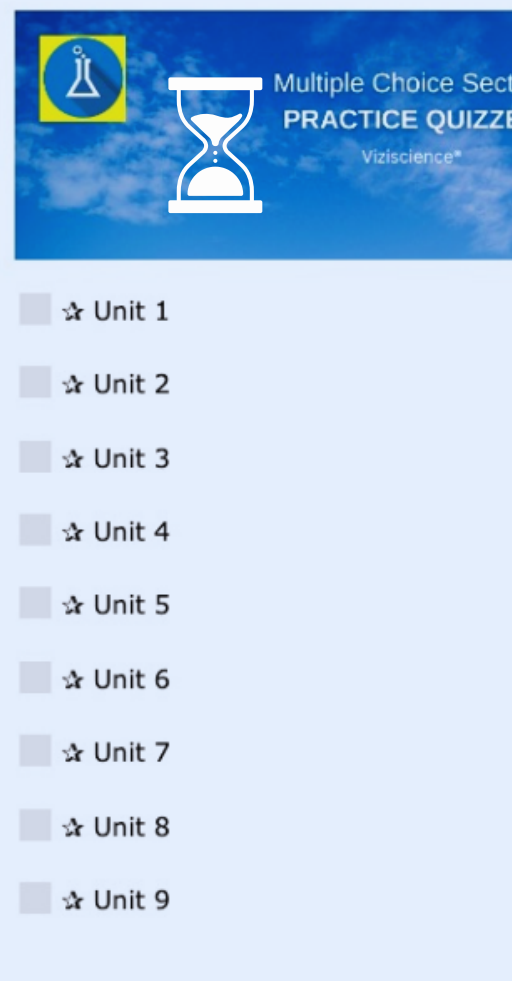
Utilize this low-stakes multiple-choice quiz as a valuable tool to help students hone their test-taking skills and enhance their confidence.



PRACTICE IS KING

The quizzes are grouped by units (Units 1 through 9), each containing 15 multiple-choice questions. These 45-minute quizzes work for both in class and at home. Teachers coordinate the start with a password release for simultaneous access.

Teachers have access to the answer key and can address concerns the following day, offering valuable insights and strengthening students' knowledge.



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GET AHEAD OF THE CURVE



The Summer Prep course includes fundamental topics like atomic structure, chemical formulas, equations, and stoichiometry.



For educators, this course is a valuable tool to gather information about incoming students, allowing them to **identify students' strengths and weaknesses**. This knowledge enables teachers to plan ahead for the class effectively.

Students benefit by gaining a head start, potentially saving up to two weeks of instructional time during the academic year to direct on critical areas.



Interactive Modules

Now that you've learned about the structure of our course, let's delve into a mini lesson and see how it works

Unit 3.1(i) - Intermolecular Forces.



This is the largest Unit in AP chemistry, it carries up to 22% exam weighting, so we want to make sure students are well prepared for this lesson.

■ 2.6 Bond Hybridization
UNIT 3
3. Intermolecular Forces and Properties
■ 3.1 (i) Intermolecular Forces (IMF)
■ 3.1 (ii) IMF, Solubility & Visual Representations
■ ✎ Intermolecular Forces
■ 3.2 Avogadro's Law & Dalton's Law
■ 3.3 Kinetic Molecular Theory
■ 3.4 Ideal Gas Law & Deviation
■ 3.5 Electromagnetic Radiation



Unit 3 is a big leap for students!



While Unit 3 doesn't require extensive mathematical calculations, students often struggle with a multitude of new concepts, unfamiliar terminology, and the need to swiftly apply their knowledge. This can lead to a loss of confidence as they transition into tackling complex questions.

One effective approach having students use our units for preparation, they can gradually build a strong foundation. This approach helps them feel well-prepared and confident when they enter your class, so they won't feel like 'fish out of water'.

USE VIZISCIENCE TO BOOST STUDENTS' CONFIDENCE

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3.1 (i) Intermolecular Forces (IMF)

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Common struggles

Students often find the following topics challenging in Unit 3:

- 1. Molecular Polarity:** Recognizing molecular polarity and its impact on intermolecular forces can be confusing for students.
- 2. Predicting Physical Properties:** Students may have difficulty predicting and explaining how intermolecular forces affect physical properties like boiling point, melting point, and solubility.
- 3. Common Misconceptions:** Some students may mistakenly believe that intermolecular forces are the same as chemical bonds or that all molecules have hydrogen bonding.

The screenshot shows a web page from Viziscience.com. The page title is "3.1 (i) Intermolecular Forces (IMF)" and the sub-heading is "Surface Tension". Below the heading is a photograph of a leaf with water droplets. The text below the photo reads: "Surface tension is responsible for the shapes of water droplets". Below this, there is a paragraph: "Water droplets are made by surface tension. Intermolecular forces are acting in all directions. At the surface, the forces from above so they are pulled together, called **surface tension**." The footer of the page includes "Privacy Policy | Terms", "Copyright © Viziscience® 2019. All rights reserved.", "Disclaimer: AP® is a trademark owned by the College Board, which is not affiliated with", and "Viziscience.com".



What essential topics need to be reviewed?

In this unit on intermolecular forces (IMF), we've included a structured review to help students revisit fundamental concepts, including electronegativity, distinguishing between ionic and covalent compounds, and understanding molecular polarity.

We will conduct comprehensive exercises to help students in identifying polar and non-polar molecules, as this is a specific area of challenge. This foundational review is crucial for mastering Unit 3.

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3.1 (i) Intermolecular Forces (IMF)

Some examples of polar and non-polar molecules:

non-polar	polar
non-polar	polar
non-polar	non-polar

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What essential topics need to be reviewed?

Students often find it challenging to differentiate between different types of intermolecular forces, including London dispersion forces, dipole-dipole interactions, and hydrogen bonding.

By providing students with numerous examples, illustrations, simulations, and videos, our goal is to ensure that, by the end of the lesson, they will have a comprehensive understanding of various intermolecular forces, their mechanisms, prevalence in different substances, and the ability to rank molecules based on the relative strengths of these forces and factors affecting them.

The screenshot shows a website interface for 'Vizscience® Interactive Modules'. The page is titled '3.1 Intermolecular forces' and is divided into two columns. The left column is titled 'Intermolecular forces' and lists four types: London dispersion forces, Hydrogen bonding, Dipole-dipole interaction, and Induced dipole-dipole interaction. The right column is titled 'Ionic compounds' and lists two types: Ion-ion attraction and Electrostatic attraction. Below the table, there is a section for 'Attraction between molecules' and 'Attraction between ions'. The website footer includes 'Privacy Policy | Terms', 'Copyright © Vizscience® 2019. All rights reserved.', and 'Vizscience.com'.

Intermolecular forces	Ionic compounds
<ul style="list-style-type: none">The London dispersion forcesHydrogen bondingDipole-dipole interactionInduced dipole-dipole interaction	<ul style="list-style-type: none">Ion-ion attractionElectrostatic attraction

Attraction between molecules

Attraction between ions.

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Let's now explore the powerful features of each page in this lesson activity and discover how they support your students' learning.





CONCEPT VIDEOS

Chemistry videos – we're well aware that they can be long and boring, losing students' interest. That's precisely why we've crafted our signature concept videos.

These videos provide in-depth exploration of each topic while maintaining a focused and clear trajectory.

Unlike YouTube videos, ours are exceptionally concise and notably free from advertisements which can disrupt valuable study time.



Teachers love our Viziscience concept videos, and they've told us how much these videos help students quickly understand tricky concepts and clear up any misconceptions.

The screenshot displays the Viziscience website interface. At the top, the navigation bar includes 'My courses (1)', 'email: support@viziscience.com', and a 'LIFESKIP' button. The breadcrumb trail reads: 'Dashboard > Courses > Advanced Chemistry > Viziscience® Interactive Modules > UNIT 3 > 3.1 (1) Intermolecular Forces (IMF) > Preview'. Below the navigation are buttons for 'Site home', 'Dashboard', and 'All courses'. The main content area features a video player titled 'Introduction to Intermolecular Forces' with a play button and a 'Flag question' link. The video shows a 3D model of water molecules with the text 'Intermolecular forces' overlaid. Below the video, a section titled 'Intermolecular Forces Vs Intramolecular Forces' explains the difference. It defines intramolecular forces as forces between atoms in a molecule, illustrated by a diagram of two overlapping circles with a red arrow pointing to the shared area. It then defines intermolecular forces as forces between molecules, illustrated by a diagram of two pairs of overlapping circles with a red double-headed arrow between them. The footer contains 'Privacy Policy | Terms', 'Copyright © Viziscience® 2019. All rights reserved.', a disclaimer about AP® being a trademark of the College Board, and the website URL 'Viziscience.com'.



SIMULATIONS

While simulations are invaluable tools to offer students a dynamic and engaging experience in chemistry, overly complex simulations will hinder learning.

What sets Viziscience simulations apart is their **user-friendliness**—students can dive right in without the need to spend valuable time learning how to operate the software.

Polarizability

What is polarizability?

Polarizability is a measure of how easily an electron cloud is distorted by an electric field. Polarizability allows us to better understand how **non-polar** atoms and molecules interact with other electrically charged atoms or molecules.

A non-polar substance (eg. carbon dioxide or methane) has evenly distributed electrons in their electron clouds. If the substance is subjected to an electric field, the electron cloud will distort.

Larger atom is more polarizable - easier to distort the electron cloud.

Smaller atom is less polarizable - harder to distort the electron cloud.

Larger atoms and molecules are **easier to polarize** than smaller ones. Why do you think smaller atoms or molecules are harder to polarize?

[CLICK FOR ANSWER](#)

Capillary Action

Capillary action is the spontaneous flow of a liquid in a narrow tube against gravity. For example, water or ethanol in a very thin tube can rise against gravity without being pumped. This movement of a liquid is a result of the forces of **cohesion** and **surface tension**.

Water climbs up a thin glass vessel against gravity and rise above the level of water in the glass. This is called capillary action.

Water is more attracted to the glass than to each other. Cohesion between water and the side of the vessel, together with surface tension, creates an upward force on the water.



REAL-TIME SKILL PRACTICE

Viziscience uses small concept check quizzes throughout the lessons, allowing students to practice their skills in real-time and assess their comprehension.

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Question 13
Tries remaining: 2
Flag question

Concept check

Propane (boils at -42°C) C ₃ H ₈ Non-polar	Dimethyl ether (boils at -24°C) CH ₃ OCH ₃ Polar	Ethanol (boils at 78°C) CH ₃ CH ₂ -O-H Polar

Propane and dimethyl ether are organic molecules. Ethanol is an alcohol. Which ones exist as gas at room temperature? Room temperature is usually considered to be about 25°C.

Propane
 Dimethyl ether
 Ethanol

Ethanol (C₂H₅OH) is a polar molecule that has a very high boiling point of 78°C. Predict the type of intermolecular forces that is most responsible for the high boiling point of ethanol.

1. London-dispersion
2. dipole-dipole
3. hydrogen bonding

Answer:

Comparing molecular mass and dipole moment of the five compounds listed in the table below, which compounds should have the highest boiling point?

Compound	Molecular mass (amu)	Dipole moment (D)
(1) Propane, CH ₃ CH ₂ CH ₃	44	0.1
(2) Dimethylether, CH ₃ OCH ₃	46	1.3
(3) Methylchloride, CH ₃ Cl	50	1.9
(4) Acetaldehyde, CH ₃ CHO	44	2.7
(5) Acetonitrile, CH ₃ CN	41	3.9

Answer:

CHECK

PREVIOUS PAGE | NEXT PAGE

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STUDENT MOTIVATION

Generally when homework is not graded, students won't complete them.

Consequently, teachers are compelled to use valuable class time to teach fundamental concepts.

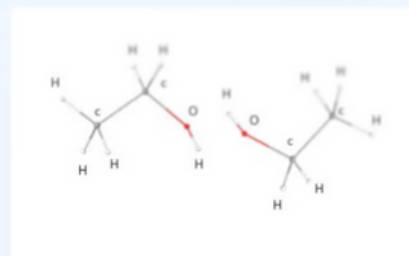
Viziscience offers an **autograding system**, which empowers teachers to seamlessly assign homework and offer **immediate feedback** to students, fostering their confidence in completing assignments and **developing self-study skills** which is crucial for AP chemistry success.



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Site home Dashboard All courses
Question 13
Tries remaining: 2
Flag question
Concept check

Answer:
Propan-1-ol (boiling point 35.5°C)
Propanoic acid (boiling point 141°C) is a gas at room temperature.
Ethanol (boiling point 78°C) and transition 1 is a gas at room temperature.
Ethanol (boiling point 78°C) is a liquid.

Ethanol (C_2H_5OH) has a relatively high boiling point because of hydrogen bonding (picture below)



(Note: Dipole moments tell us about the charge separation in a molecule. A molecule with a very large dipole moment is very polar and has strong intermolecular forces.)

(Compound 3) Acetonitrile, CH_3CN should have the highest boiling point because its dipole moment is the largest. The mass is irrelevant for comparing boiling points because all three molecules have about the same molecular mass.



VALUABLE REPORTS FOR TEACHERS

First name / Surname	Email address	State	Started on	Completed	Time taken	Grade/10.00	Q. 1	Q. 2	Q. 3	Q. 4	Q. 5	Q. 6	Q. 7	Q. 8	Q. 9	Q. 10
<input type="checkbox"/> Review attempt	2t...@...t	Finished	May 5 2021 9:06 AM	May 5 2021 10:40 AM	1 hour 34 mins	9.39	✓ 0.67	✓ 0.67	✓ 0.67	✓ 0.67	✓ 0.67	✓ 0.67	✓ 0.67	✓ 0.67	✓ 0.67	✓ 0.67
<input type="checkbox"/> Review attempt	4t...@...t	Finished	May 5 2021 10:20 AM	May 5 2021 11:15 AM	54 mins 30 secs	9.67	✓ 0.67	✓ 0.67	✓ 0.67	✓ 0.56	✓ 0.67	✓ 0.67	✓ 0.67	✓ 0.67	✓ 0.67	✓ 0.67
<input type="checkbox"/> Review attempt	4t...@...t	Finished	May 5 2021 10:32 AM	May 5 2021 8:09 PM	9 hours 37 mins	9.06	✓ 0.67	✓ 0.67	✓ 0.67	✓ 0.67	✓ 0.67	✓ 0.67	✓ 0.67	✓ 0.67	✓ 0.67	✓ 0.67
<input type="checkbox"/> Review attempt	4t...@...t	Finished	May 5 2021 11:42 AM	May 5 2021 2:12 PM	2 hours 30 mins	9.50	✓ 0.67	✓ 0.67	✓ 0.67	✓ 0.67	✓ 0.67	✓ 0.67	✓ 0.67	✓ 0.67	✓ 0.67	✓ 0.67
<input type="checkbox"/> Review attempt	4t...@...t	Finished	May 5 2021 11:48 AM	May 5 2021 12:56 PM	1 hour 8 mins	9.39	✓ 0.67	✓ 0.67	✓ 0.67	✓ 0.67	✓ 0.67	✓ 0.67	✓ 0.67	✓ 0.67	✓ 0.67	✓ 0.67
<input type="checkbox"/> Review attempt	2t...@...t	Finished	May 5 2021 11:51 AM	May 13 2021 11:50 AM	7 days 23 hours	9.39	✓ 0.67	✓ 0.67	✓ 0.67	✓ 0.67	✓ 0.67	✓ 0.67	✓ 0.67	✓ 0.67	✓ 0.67	✓ 0.67
<input type="checkbox"/> Review attempt	j...@...t	Finished	May 25 2021 11:58 PM	June 9 2021 12:04 PM	14 days 12 hours	8.06	✓ 0.67	✓ 0.67	✓ 0.67	✓ 0.56	✓ 0.67	✓ 0.67	✓ 0.67	✓ 0.22	✓ 0.67	✓ 0.67
Overall average						8.30	0.65	0.58	0.57	0.59	0.53	0.60	0.52	0.59	0.57	0.66



Too often, we've seen teachers feeling disheartened when they receive disappointing results after a quiz or exam. Our goal is to prevent such surprises and equip educators with the tools they need to stay ahead. The progress report will help teachers tremendously to plan ahead and offer support when needed.

Get started with Viziscience®

1. Teacher subscribe and create an account on the system.
2. Students create their own individual accounts and enroll into their respective classes assigned to their teacher.
3. Students use the activities under teacher's instructions.
4. Teachers can view students' activity reports at any time.



We hope you've found our approach to preparing students for success in AP Chemistry valuable. It's important to mention that our resources are cost-effective, offering a wealth of content and an accessible learning management system at an affordable price. Our commitment is to support educators in their mission. By subscribing, you become a part of our community, and we're here to provide you with the best resources and support possible.

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\$55 for your entire class



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(Normally \$75 - \$150)

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